

ACCESSION NR. AT4033626

Work with the machine takes place in the form of a dialogue, in which the machine evaluates the assimilation of the material and the ability to apply it to the solution of problems. It is necessary when questions arise the student assimilated poorly and the machine provides explanations are provided for these areas. The logic of the machine operation makes it possible to teach using the ramified programs shown in Figure 1, a and 1, b of the enclosure. The program shown in Figure 1, a is written on the assumption that the student is somewhat familiar with the given subject. The teaching process takes place as follows: the machine proposes to the student, sequentially one after the other, a series of not more than 10 logically connected questions on a given subject 1.0, 2.0,, N.0. Analyzing the answers of the student, the machine supplies him either with the next question over channels "f", if the answer was correct, or furnishes, over channels "a", information leading to the correct answer (hints) 1.1, 2.1, 3.1,, N.1, if the answer was incorrect. In the event of correct answers after the suggestions, the next question is proposed over channels "d". If the student answers incorrectly even after the additional information, then the machine provides exhaustive clarifications on the troublesome questions over channels "b", after which the next question is submitted over channels "e". If the stu-

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dent has difficulty in answering any of the proposed questions, he can request "assistance", in which event full information 1.2, 2.2,N.2 is immediately provided over channels "c". Operation with the program shown in Figure 1, differs from the above only in that the machine first submits the new material 1.2, 1.1, 1.2, with question 2.2 then posed as a development of what has gone before, and, if the student answers are "iv", the exposition of the subsequent material begins. If the answer is incorrect, then leading advice 2.1 and exhaustive explanations 2.2 are given, followed thereafter by the exposition of the new section 3.0, 3.1, 3.2, etc. Depending on its volume, the material can be arranged on a larger number of frames as well, but the number of frames must be a multiple of three. The material is particularly suitable to the study of individual aspects of the theory of automatic control and foreign languages. The complete OM-4 machine contains a modernized MFT-10 slide projector, control panel and set of programs with answer cards. The programs are based on motion picture film and stored on special cassettes. The length of the programmed tapes depends on the capacity of the cassette, the projector and must not exceed 54 standard film frames. The sequence with which the program texts are arranged on the film frames is indicated in appendices to the article. The material is projected on a screen and the answers are introduced

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by the circuit is set by means of buttons. For each new value values are set on these buttons and, simultaneously with the of teaching machine, correct and incorrect answer card sets which are interchangeable together with the film-program. The circuit the circuit is described, certain particularities of the program machine are discussed and certain conclusions and possible improvements advanced. Two examples of programming are given in the appendix: the article one from the theory of automatic control, the other dealing with the translation of a complex German sentence into Russian. Orig. art. has 3 figures.

ABSTRACTION: none

SUBMITTED: 03Dec63

SUB CODE: DP

DATE ACQ: 16Apr64

NO REF SOV: 003

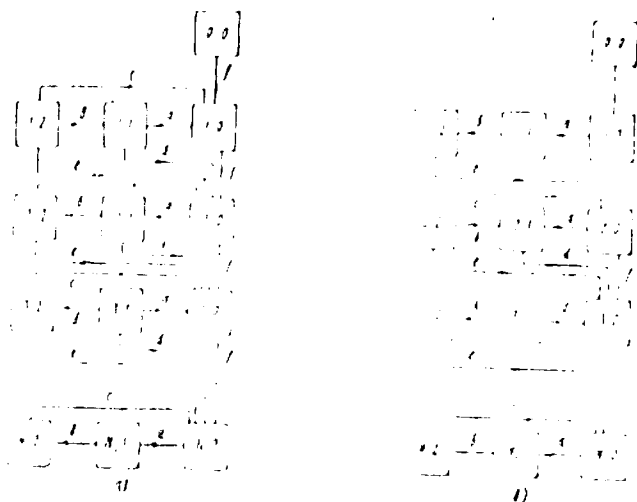
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OTHER: 001

Card 4/5

ACCESSION NR: ~~AP~~4033626

ENCLOSURE: 01



and 5/5

Figure 1. Operating logic of the CM-9 machine

SOKOLINSKIY, I.Ya.; MISHCHENKO, I.M.

Teaching machine with a ramified program. Izv. vys. ucheb.
zav.; radiotekh. 6 no.4:425-434. 71-Ag '63. (MIRA 1:11)

MISHCHENKO, I.N., Geroy Sotisialisticheskogo Truda

Communism is our cherished, vital cause. Transp. stroi. 12 no.2:5-6
F '62. (MIRA 15:7)

1. Nachal'nik smeny stroitel'stva No.2 Gosudarstvennogo stroitel'nogo
upravleniya Kiyevskogo metropolitena.
(Kiev—Subways)

CH 115

The effect of Röntgen rays on the appearance of complement-fixing bodies in the blood. I. P. Mishchenko and M. M. Fomenko. *Ann. Roentgenol. radion.* U.S.S.R. 13, 427 (1964) (English 337). Blood taken from the exptl. animals was examined by the Bordet-Gengou test with protein antigens, which were obtained from the organs (liver, kidneys, muscles) of normal animals. In the exptl. animals the same organs were x-rayed with doses varying from 20 to 100% of an H.F.D. Blood tests were made immediately after radiation, and within the next few days. The numerous observations showed regular appearance of complement-fixing bodies in the blood subsequent to the application of x-rays. This is regarded as a reaction to the appearance of foreign proteins in the blood caused by radiation. The expts. have shown that 60% of an H.F.D. is the optimal dose for causing the appearance of anti-bodies in the blood of dogs. This fact may prove of importance in studying the problem of immunity.

N. N. Menshik

ASB 3.4. METALLURGICAL TREATMENT CLASSIFICATION

Data obtained by chemical tests of the inflammatory focus subjected to roentgenotherapy. I. P. Mikhlenko and S. N. Ledanov. *Ann. roentgen. radiol.* 1955, No. 5, 322-3. Healthy rabbit skin contains 0.6 mg % amino N. When inflamed the skin contains slightly higher amt., which, however, is greatly increased when subjected to x-rays. Thus on the second day the inflamed skin contains 24.3 mg % amino N, without irradiation, and 40.0 mg % when treated with x-rays on the third day 71.4 and 107.0 mg % resp. On the following days the amino N content falls more rapidly in the irradiated tissue. An increase is also noticed in the amt. of trypsin and proteases, especially globulins. H. Cohen.

CIA-RDP86-00513R001134620007-7"

ca

The condition of hunger in cancer. I. P. Mishchenko
 Vopr. med. i. R. N. S. 2, No. 1, 104 (1977). *Chim.*
Zentr. 1979, U, 2243. Mice having Ehrlich carcinoma
 rats with Jensen and with Krichewski-Semel'nikov sar-
 coma or Flexner-Jobling carcinoma, hens with spon-
 sarcoma, rabbits with fat tumor and dogs with spon-
 sarcoma were kept in a condition of slow hunger for
 various periods, only water being given. As compared
 varying periods, all these animals showed a
 to normally hungry animals, all these animals showed a
 marked shortening of life, death occurred with a slight
 loss in wt. and growth of the tumor continued during
 hunger. The amt. of N excreted in the urine (referred to
 the body wt.) remained approx. const. and corresponded
 to the normal values. Only in the case of one batch with
 vaginal carcinoma was an excretion of twice the normal
 amt. of N observed. From the 10th day on, the urra
 fraction of the total N was sharply decreased as compared
 with the value for normal hunger. The carcinomatous
 animals excreted a definite portion of the N in the form
 of non-endoxidized products. M. G. Mironov

MISHCHENKO, I. P., Prof.

Troitsk Veterinary Inst.

"Skin-tail specimen for functional examination of the reticulo-
endothelial system in domestic animals."

SO: Veterinariya 27(3), 1950, p. 45

MISHCHENKO, I.P.

Effect of ultraviolet irradiation in the presence of argon and nitrogen
on absorption spectrum of amino acids. Biokhimiia 19 no.3:268-72.
My-Je '54. (MLRA 7:8)

(AMINO ACIDS, effect of radiations on,
ultraviolet rays, absorp. spectrum after irradiation in
presence of argon & nitrogen)

(ULTRAVIOLET RAYS, effects,
on amino acids absorp. spectrum, irradiation in presence
of argon & nitrogen)

(NITROGEN,
eff. of ultraviolet irradiation in presence of nitrogen
& argon on amino acids absorp. spectrum)

(ARGON,
eff. of ultraviolet irradiation in presence of nitrogen
& argon on amino acids absorp. spectrum)

11A

Action of ultraviolet rays on the absorption spectra of amino acids. I. P. Mishchenko, Izvestiya Akad. Nauk SSSR, 1962, No. 10, 1032-1037. The effect of changes in the absorption spectra of amino acids (20% solutions) were observed after illumination with ultraviolet rays. A 2% solution of cysteine, when exposed for several hrs., turned yellow, formed a ppt., and evolved H₂S. Cysteine also formed H₂S after only a few min. irradiation. Curves are given for changes in the absorption spectra of glycine, alanine, serine, cysteine, cystine, taurine, leucine, lysine, asparagine, arginine, tyrosine, tryptophane, and histidine. H. P. Swick.

MISHCHENKO, I.P.

Erythrocyte weight coefficient as an indicator of oxidation
intensity in animals. Izv.Otd.est.nauk AN Tadsh.SSR no.9:
89-100 '55. (MLRA 9:10)

1. Tadzhikskaya nauchno-issledovatel'skaya veterinarnaya opyt'naya
stantsiya.

(OXIDATION, PHYSIOLOGICAL) (ERYTHROCYTES)

MISHCHENKO, I.P.

Cellular reaction of skin of animals to irritation. Izv.Otd.
est.nauk AN Tadzh.SSR no.9:101-108 '55. (MLRA 9:10)

1. Tadzhikskaya nauchno-issledovatel'skaya veterinarnaya
opytnaya stantsiya.
(Irritation)

KNIGA, M.V.; MISHCHENKO, K.P.

Ice calorimeter with electrical calibration. Zhur. prikl. khim. 30
no.11:1708-1711 N '57. (MIRA 11:2)

1. Institut khimii AN Estonskoy SSR.
(Calorimeters)

KNIGA, M.V.; VASIL'YEVA, T.M.; MISHCHENKO, K.P.

Possibility for evaluating the specific surface of kukersite shale
on the basis of its heat of interaction with liquids. Zhur.prikl.
khim. 30 no.12:1866-1868 D '67. (MIRA 11:1)

1. Institut khimii Estonskoy SSR.
(Oil shales)

MISCHENKO, K.P.; POMINOV, I.S.

Solvation of ions in electrolyte solutions. Part 5: Coordination numbers of copper, cobalt, and neodymium ions associated with their hydration in alcohol aqueous solutions [with summary in English]. Zhur.fiz.khim. 31 no.9:2026-2031 S '57. (MIRA 11:1)

1. Leningradskiy tekhnologicheskii institut. Kazanskiy gosudarstvennyy universitet im. V.I. Ul'yanova-Lenina.
(Solvation) (Copper) (Cobalt) (Neodymium)

1. The first step in the investigation of the effect of the position of the body on the rate of heat loss is to determine the rate of heat loss in the supine position.

2. The second step is to determine the rate of heat loss in the sitting position. This is done by having the subject sit on a chair and measure the rate of heat loss.

3. The third step is to determine the rate of heat loss in the standing position. This is done by having the subject stand and measure the rate of heat loss.

4. The fourth step is to determine the rate of heat loss in the walking position. This is done by having the subject walk on a treadmill and measure the rate of heat loss.

5. The fifth step is to determine the rate of heat loss in the running position. This is done by having the subject run on a treadmill and measure the rate of heat loss.

and 6. The sixth step is to determine the rate of heat loss in the swimming position. This is done by having the subject swim in a pool and measure the rate of heat loss.

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the Calculation of the Fundamental Thermodynamic Values for HClO_4 at 25°C Centigrade

SVT/70-100-11/48

$\Delta H_f^\circ(\text{HClO}_4) = -103.1 \text{ kJ/mol}$

$S^\circ(\text{HClO}_4) = 107.5 \text{ J/mol}^\circ\text{K}$

$S^\circ(\text{HClO}_4) = 11.5 \text{ J/mol}^\circ\text{K}$

where ΔH_f° is the standard enthalpy of formation, S° is the standard entropy.

S. I. Zaitsev June 1, 1977

SVT/70/

MISHCHENKO, K.P., prof.; BOMSHTEYN, T.B., dotsent

[Problems of higher chemical and technical education; symposium papers presented at the Eighth Mendeleev Congress on General and Applied Chemistry] Problemy vysshego khimicheskogo i tekhnologicheskogo obrazovaniia; doklady na simpoziume na VIII Mendelevskom s"ezde po obshchei i prikladnoi khimii. Moskva, 1959. 54 p. (MIRA 12:10)

1. Leningradskiy tekhnologicheskii institut tsellyulozno-bumazhnoy promyshlennosti (for Mishchenko). 2. Moskovskiy institut khimicheskogo mashinostroyeniya (for Bomshteyn). (Chemistry--Congresses)

MISHCHENKO, K. P.

24(8) PLENARY REPORTS NOV/28/69
Akademiya nauk SSSR. Otdeleniye khimicheskikh nauk
Termodinamika i stroeniye rastvorov, teoriya soveshchaniya...
(Thermodynamics and Structure of Solutions, Transactions of the
Conference Held January 27-30, 1968) Moscow, Izd-vo AN SSSR,
1979. 295 p. 3,000 copies printed.

Ed.: R. I. Sharapov, Doctor of Chemical Sciences, Ed. of Publishing
House: N. O. Veprev, Tech. Ed.: T. V. Polyakova.

PURPOSE: This book is intended for physicists, chemists, and
chemical engineers.

CONTENTS: This collection of papers was originally presented at the
Conference on Thermodynamics and Structure of Solutions sponsored
by the Academy of Chemical Sciences of the Academy of Sciences,
USSR, and the Department of Chemistry of Moscow State University,
and held in Moscow on January 27-30, 1968. Officers of the
conference are listed in the Foreword. A list of other reports
concerning the problems, but not included in this book,
are given. Among the problems treated in this book are
electrolytic solutions, ultrasonic measurement, viscosity,
and thermodynamic properties of various mixtures, specific
scopic analysis, etc. References accompany individual papers.

Thermodynamic Properties of Water in Solutions of Electrolytes

Antonova, L. G., and M. I. Sharapov. Verification of the
Theory of Molecular Dispersion of Light by Means of Binary
Solutions

Vukobratovic, R. P. Anisotropic Dispersion of Light and Its Use in
Studying Liquids and Solutions

Chernomir, L. P., and A. M. Ponomarev. Partial Molar
Entropies in Systems Acetic Acid-Water and Phenol-Acetic
Water and the Structure of These Solutions
Chernomir, L. P., and A. M. Ponomarev. Spectroscopic Methods for Studying the
Structure of Solutions

Belitsky, M. G. Spectroscopic Methods for Studying Complexes in
Solution

Belitsky, M. G., V. P. Kabanov, and L. I. Kabanov.
Relaxation Processes in the Absorption Spectra of
Radiation of Solutions of Organic Substances and the Chemical
Nature of Solvents

Vaynshteyn, L. Ya., and L. I. Kabanov. Spectroscopic Study of
Solvation of Ions in Solutions with the Aid of Solvents.
Absorption Spectra

Antikarov, I. L. Study of the Effect of the
Solvation of Ions on the State of the Thermodynamic Values
of Absorption Spectra of Solutions and Acid-Base
Equilibria. V. P. Chernomir, and M. V. Chernomir.
Infrared Spectra of Electrolytic Solutions in Acetic Acid

Levin, V. L., Ye. O. Ponomarev, L. O. Dzhuravskaya, and
L. V. Levshina. Study of Association in Chemical
Solutions of Dyes by Means of Absorption and Luminescence
Spectra

Levin, V. L., V. O. Ponomarev, L. O. Dzhuravskaya, and
L. V. Levshina. Effect of Ionization and Association on
Optical Properties of Complex Organic Molecules

MENDELEYEV, Dmitriy Ivanovich; ~~MISHCHENKO~~, K.P., prof., red.; ZAYCHIK,
N.K., red. izd-va; PEVZNER, R.S., tekhn. red.

[Solutions] Rastvory. Red. i stat'ia K.P. Mianchenko. Izd-vo
Akad. nauk SSSR, 1959. 116 p. (MIRA 1:4)
(Solution (Chemistry))

YAKOV V. V. (1911-1981) (1911-1981) (1911-1981)

"The first of the three main types of the
Yakov V. V. (1911-1981) (1911-1981) (1911-1981)
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MISHCHENKO, K.P.; TALUD, S.L.; YAKIMOVA, V.I.

Reaction of cellulose with liquids. Vyssokospeed. 1 no. 5:
662-669 My '69. (MIRA 12:10)

1. Leningradskiy tekhnologicheskii institut tsellulozno-bumazhnoy
promyshlennosti.
(Cellulose) (Thermochemistry)

On the Dissociation of Arsenic Acid

SCV 78-4-2-5 40

were determined. The results are shown in figures 2 and 3. The heat effects, the change of the isobaric potential, and the change of the entropy in the processes 1, 2, and 3 of the gradual dissociation were determined. The average values of the thermodynamic quantities of the dissociation processes of H_3AsO_4 in aqueous solutions are shown in table 4.

There are 3 figures, 4 tables, and 35 references, 6 of which are Soviet.

ASSOCIATION: Leningradskiy nauchno-issledovatel'skiy institut pererabotki nefti (Leningrad Scientific Research Institute of Petroleum Processing); Leningradskiy tekhnologicheskii institut (Leningrad Technological Institute);

SUBMITTED: November 5, 1967

Card 2/2

MISHCHENKO, K.P., doktor khimicheskikh nauk

Thermodynamics of some processes of woodpulp and paper
production. [Trudy] NTO bum.i der.prom. no.8:52-53 '59.
(MIRA 16:2)

(Woodpulp industry--Research)

5(4)

1977 - 1 - 10

AUTHORS: Mishchenko, K.I., Talmuz, L.L. and Yakimov, V.I.

TITLE: On the Value of the Specific Surface of Cellulose

PERIODICAL: Kolloidnyi Zhurnal, 1977, Vol. 39, No. 1, p. 10-11 (USSR)

ABSTRACT: The present investigation is concerned with the determination of reliable methods for determining the determination of the value of the specific surface of cellulose in the dry and in swollen states. The author also tried to determine the most probable value of the specific surface of standard cotton cellulose and technical wood cellulose obtained by different methods. For the determination of the specific surface of cellulose in the dry state, the standard method of nitrogen vapor adsorption at its boiling point (-195.8°C) was used. For the determination of the specific surface of cellulose in the swollen state the method of ion exchange, proposed by V.I. Yakimov, was used. As must be pointed out, the author of the present investigation confirmed the reliability of the reliable method.

Car: 1/1

On the Value of the Specific Surface of Cellulose

The specific surface of standard cotton cellulose was found to be 16-18 m²/g. The results obtained with the two methods are in agreement. It was further stated that steaming in water does not have an effect on the value of the specific surface of the plant fiber. Wood celluloses obtained with various methods are sharply distinguished from natural fiber, as far as their internal structure is concerned. The specific surface of wood cellulose was found to be 10-12 m²/g, of cellulose in the form of paper - 100-150 m²/g. The authors mention the names of A.V. Kiselev, T. Pikkerstaf, V.I. Kuznetsov, and N.I. Nikitin. There are 10 references, 10 of which are English, Soviet and German.

ASSOCIATION: Leningradskiy tekhn. i inzh. inst. (LIT) -
Card 2/3 bumazhnyy promyshlennost', Kafeira fizicheskoy

On the Value of the Specific Unit: ...

kollektivnyy kharakter (collective character) of the ...
of the ... and ...
and ...

SUBMITTED: 2 Oct 1971

Carl ...

5(4)

AUTHORS: Mishchenko, K. P., Yakovlev, I. P. SOV. 79-29-11172

TITLE Thermochemistry of Electrolyte Solutions (Termokhimiya rastvorov elektrolitov) V. Integral Heats of Solution of NaCl and $MgCl_2 \cdot 6H_2O$ in Water at $+t^\circ$ and $-t^\circ$ (V. Integral'nyye teploty rastvoreniya $NaCl$ i $MgCl_2 \cdot 6H_2O$ v vode pri temperaturakh $+t^\circ$ i $-t^\circ$)

PERIODICAL: Zhurnal obshchey khimii, 1979, Vol. 29, No. 4, pp. 1771-1774 (USSR)

ABSTRACT: The thermochemical investigations of aqueous solutions of electrolytes in a large number of different concentrations and temperatures is of considerable interest both with respect to the evolution of the theory of the solutions and for some industrial fields. The polytherms to be obtained of the integral heats of solution, the heat capacities and partial products (molar ratio) of these values make it possible, in connection with other constants, to characterize more thoroughly the nature of the electrolyte solutions and to clarify especially the part the solvent plays in the formation of the solution. In aqueous solutions this part

Card 1/3

Thermochemistry of Electrolyte Solutions

V. Integral Heats of Solution of NaCl and $MgCl_2 \cdot 6H_2O$ in Water at $+2$ and -2

had to become particularly evident at temperatures above and below 0° since in this case the thermal motion less interferes with the primary water structure. Thus the interactions between the ions and the solvent are becoming manifest in a considerably pronounced way under these conditions. The integral heats of solution of NaCl, $MgCl_2 \cdot 6H_2O$ and $MgCl_2$ were determined at $+2$ and -2° in a wide concentration range. The corresponding partial enthalpies of water and electrolyte were calculated in the molar ratio under equal conditions. The comparison of the dependences of these properties on concentration and temperature with the data available in publications permitted some conclusions concerning the structure of the concentrated solutions and their change at different temperatures. There are 4 figures, 5 tables, and 10 references 11 of which are Soviet.

Card 2/3

Thermochemistry of Electrolyte Solutions.

0 7,79-29-11 71

V. Integral Heats of Solution of NaCl and $MgCl_2 \cdot 6H_2O$ in Water at +2 and -62

ASSOCIATION: Leningradskiy tekhnologicheskii institut imeni Lensovet
(Leningrad, Technological Institute imeni Lensovet)

SUBMITTED: June 9, 1958

Card 3/3

"APPROVED FOR RELEASE: 06/14/2000

CIA-RDP86-00513R001134620007-7

APPROVED FOR RELEASE: 06/14/2000

CIA-RDP86-00513R001134620007-7"

TITLE: ANALYSIS OF THE EFFECTS OF VARIOUS TEMPERATURES
 ON THE BLEACHING OF COTTON FABRICS
 ABSTRACT: The effects of various temperatures on the bleaching of cotton fabrics were studied. The results show that the bleaching rate increases with increasing temperature. The highest bleaching rate was observed at 100°C. The results also show that the bleaching rate is higher for cotton fabrics than for other types of fabrics. The results of this study can be used to optimize the bleaching process for cotton fabrics.

(4)

AUTHORS.

Plis I. Ye. Mishchenko, K. P.,
Troitskaya, N. V.

SCV/76-33-8-11/32

TITLE.

Potentials of Chlorine Electrodes at Various Temperatures

PERIODICAL.

Zhurnal fizicheskoy khimii: 1959, Vol 33, Nr 8, pp 1744 - 1749
(USSR)

ABSTRACT

The oxidizability of chlorine and its compounds is important for the technology of chlorination and bleaching of cellulose and textile fabrics. In publications many investigations concerning the properties of chlorine and particularly regarding the determination of the potential (P) of the chlorine electrode (CE) are described. In (Ref 2) it was found that a platinum electrode (PE) behaves like a (CE) in acid hypochlorite solutions. On the basis of data found in publications, the (P) of the (PE) in acid hypochlorite solutions was investigated in the present case. The solutions contained larger amounts of dissolved chlorine (C). It was assumed that the values obtained were due to the balance $\frac{1}{2} \text{Cl}_{2, \text{gas}} + e \rightleftharpoons \text{Cl}_{\text{solution}}^-$ (2). Potentiometrical measurements were carried out in the most practical

Card 1/3

Potentials of Chlorine Electrodes at Various Temperatures

SOV/70 11 11/13

temperature range at 10, 25, 35, and 50°C. All (E) and pH measurements were carried out by the compensation method with a Raps potentiometer (with an electrometric tube 18C1). A series of potentiometrical titrations with a Pt and glass electrode of 0.08 - 0.04 n NaClO-solutions and 0.1 n H₂SO₄-solutions

were carried out, the pH and the oxidation potential (OP) being measured. The calculation of the normal potential of (C), $\varphi_{Cl_2/2Cl^-}$

was carried out by means of a known equation (Table 1) for solutions with a (C)-concentration corresponding to the (C) pressure in equilibrium at 1 atm). The normal (OP) of the system Cl₂ gas - 2 Cl⁻ solution for the above temperatures were calculated from the experimental data (Table 2). The values for 25°C agree well with those found in publications (Refs 5, 6, 7). It is assumed that for this reason the values given for other temperatures are also reliable. Equations for the temperature function of $\varphi_{Cl_2/2Cl^-}$ and $\Delta Z_{Cl_2/2Cl^-}^0$ (change in the isobaric

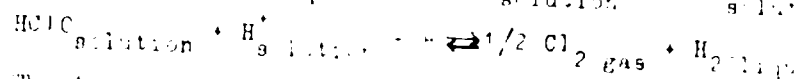
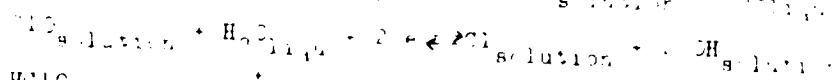
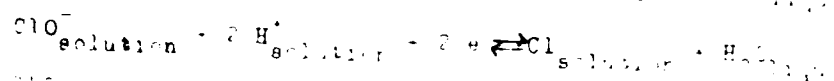
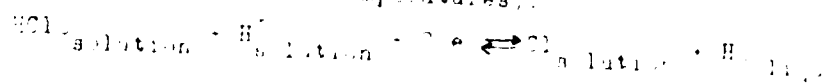
Card 2/3

potentials) were obtained and the values ΔZ^0 , $d\varphi/dT$, ΔH and ΔS

Potentials of Chlorine Electrodes at Various Temperatures

SOV, 76, 2, 2, 11

of the equilibria at the above temperatures were calculated. The values φ^0 and ΔZ were determined for the following equilibria (at the above temperatures).



The temperature functions of the normal potentials of the latter equilibria are given by corresponding equations. There are 3 tables and 23 references, 11 of which are Soviet.

ASSOCIATION. Leningradskiy tekhnicheskii institut (Leningrad Technical Institute)

SUBMITTED. January 10, 1958
Card 3/3

KISELEVA, Ye.V.; KARETNIKOV, O.S.; KUDRYASHOV, I.V.; BOTVINKIN, O.K., doktor
khim.nauk, retsenzent; MAKOLKIN, I.A., doktor tekhn.nauk, retsenzent;
MISHCHENKO, K.P., doktor khim.nauk, retsenzent; GRYAZNOV, V.M.,
red.; REZUKHINA, T.M., red.; ZAZUL'SKAYA, V.P., tekhn.red.

[Collection of illustrated physical chemistry problems and exercises]
Sbornik primerov i zadach po fizicheskoi khimii. Moskva, Gos.
nauchno-tekhn.izd-vo khim.lit-ry, 1960. 264 p. (MIRA 13:7)
(Chemistry, Physical and theoretical--Problems, exercises, etc.)

MISHCHENKO, K.P., prof.

Work of the Leningrad Administration of the D.I.Mendeleev All-
Union Chemical Society. Zhur. VKhO 5 no.4:466-468 '60.

(MIRA 13:12)

(Leningrad--Chemical societies)

MISHCHENKO, K.P.; TUMANOVA, T.A.; FLIS, I.Ye.

Determination of sulfide, hydrosulfide, and pyrosulfide
present simultaneously in pure aqueous solutions. Zhurnal khim.
fiz. no.2:21-217 Mr-Apr 1969. (MIRA 14:7)

1. Leningradskiy tekhnologicheskii institut tselulozno-papernoy
promyslovenosti.
(Sulfide) (Hydrosulfide)

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MISHCHENKO, K.P., prof., doktor khim.nauk

N.A.Ismailov's book "The Electrochemistry of Solutions." Reviewed by K.P.Mishchenko. Zhur.prikl.khim. 33 no.7:396-397
Jl '60. (MIRA 13:7)

(Electrochemistry)
(Ismailov, N.A.)

11.131°

AUTHORS: MISHKIN, I. I. (Moscow, U.S.S.R.)

TITLE: The effect of the temperature of the hydrogen peroxide decomposition on the rate of the reaction of the decomposition of the hydrogen peroxide at different temperatures.

PERIODICAL: Zhurnal Prikladnoi Khimii, 1967, No. 10, p. 1671.

TEXT: Data on the kinetics of the decomposition of the hydrogen peroxide at different temperatures are presented. The results of the study of the reaction of the decomposition of the hydrogen peroxide at different temperatures are presented. X

at temperatures of 10-40°C. The results of the study of the reaction of the decomposition of the hydrogen peroxide at different temperatures are presented. Vreyskiy, I. I. (Moscow, U.S.S.R.)

Card 1

[illegible]

MISHCHENKO, K.P.; PODGORNAYA, Ye.A.

Effective radii of water molecules in solvate ion sheaths as a
characteristic of solvate sheath orderliness. Trudy LTI no.61:
21-24 '60. (MIRA 15:5)
(Solvation) (Electrochemistry)

KLYUYEVA, M.L.; MISHCHENKO, K.P.; FEDOROV, M.K.

Solubility of picric acid in methyl alcohol in the temperature
range from 5° to 50°. Trudy LTI no.61:47-51 '60. (MIRA 15:5)
(Picric acid) (Solubility)

MISCHENKO, K. P. (Mishchenko, K. P.), BEREGL, Endre (translator)

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Thermodynamics of the interaction of cellulose with water
and water solutions of electrolytes. Report No.3. Trudy
LTA no.91:65-70 '60. (MIRA 15:12)

1. Leningradskiy tekhnologicheskoy institut tsellyulozno-
bumazhnoy promyshlennosti.
(Cellulose) (Electrolytes) (Thermodynamics)

MISHCHENKO, K.P.; PONOMAREVA, A.M.; RAVDEL', A.A.; BARON, N.M.;
YEGOROV, I.M.; KVIAT, E.I.; VOLOVA, Ye.D.; MARKOVICH, V.G.;
SEMENOV, G.I.; MARGOLIS, V.N.; SFORODINA, T.P.; YAVORSKIY,
I.V. Prinsipal uchastiye FRANK-KAMENETSKIY, V.A.; TOMARCHENKO,
S.L., red.; LEVIN, S.S., tekhn. red.

[Practical work in physical chemistry] Prakticheskie raboty po
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(Chemistry, Physical and theoretical--Laboratory manuals)

MISHCHENKO, K.P., doktor khimicheskikh nauk, ELIS, I.Ye., kand.khimich. nauk,
BYNYAYEVA, M.K., kand.khimich. nauk, KRYUKOVA, Z.M., kand.khimich.
nauk; SALNIS, K.Yu., kand.khimich. nauk; BLOSHTEYN, I.I., inzh.;
DOBRYSHIN, K.D., inzh., PISH, S.I., inzh.

Technology of the production of chlorine dioxide. Trudy LTITSHP
no.8:81-88 '61. (MIRA 1:40)

(Chlorine oxides)

MISHCHENKO, K.P., doktor khimicheskikh nauk; DYMARCHUK, N.P., kand.
khimicheskikh nauk; SOLOLOV, M.M., inzh.

Thermodynamics of the interaction of cellulose with water and
water solutions of electrolytes. Part 2: Comparing the temperature
relationships of the heat of interaction of water with standard
and viscose celluloses. Trudy LTITSBP no.8:109-113 '61.

(MIRA 10:3)

(Heat of adsorption) (Woodpulp)

MISHCHENKO, K.P., doktor khimicheskikh nauk; DYMARSHUK, V.P., kand.
khimicheskikh nauk; SHKOTINA, L.V., inzh.

Thermodynamics of the interaction of cellulose with water and
water solutions of electrolytes. Part 4: Effect of predissolved
nitrates on the integral heat of the interaction of cellulose
with water. Trudy LIT'SBP no.8:114-119 '61. (MIRA 1:1)
(Woodpulp) (Nitrates) (Heat of adsorption)

MISHCHENKO, K.P.

M.V. Lomonosov founder of the law of conservation (on the 250th anniversary of his birth). Izv. vys. ucheb. zav.; khim. i khim. tekhn.
4 no. 6: 882-888 '61. (MIRA 15:3)
(Lomonosov, Mikhail Vasil'evich, 1711-1765)

MISHCHENKO, K.P., prof.; PODGORNAYA, Ye.A.

Thermochemistry of electrolyte solutions. Part 6: Aqueous
solutions of CoCl_2 at different temperatures. *Zhur.ob.khim.* 31
no.6:1743-1745 Je '61. (MIRA 14:6)

1. Leningradskiy tekhnologicheskii institut imeni Lensovet.
(Cobalt chloride) (Heat of solution)

KUSTODINA, V.A.; MISHCHENKO, K.P.; FLIS, I.Ye.

Thermodynamic characteristics of the interaction between chlorine
monoxide and sodium hydroxide. Zhur. prikl. khim. 34 no.1:125-129
Ja '61. (MIRA 14:1)

(Chlorine oxide)

(Sodium hydroxide)

FLIS, I.Ye.; MISCHENKO, K.P.; KUSTODINA, V.A.

Thermodynamic characteristics of the reaction between chlorine
monoxide and hydrogen peroxide in an alkaline medium. Zhur. prikl.
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(Chlorine oxide) (Hydrogen peroxide)

MISHCHENKO, K.P.; FEDOROV, M.K.

Activities of the solvent in methanol solutions of picric
acid in the temperature range from 5° to 50°. Zhur.prikl.
khim. 34 no.8:1889-1891 Ag '61. (MIRA 14:8)
(Picric acid) (Methanol)

MISHCHENKO, K.P.; FEDOROV, M.K.

Structure of picric acid solutions in methanol. Zhur.strukt.khim.
3 no.1:15-20 Ja-F '62. (MIRA 15:3)

1. Leningradskiy tekhnologicheskij institut imeni Lensoвета.
(Picric acid) (Methanol)

LYMARCHUK, N.P., kand. khim. nauk; KULRYAVTSEVA, I.V., inzh.;
MISHCHENKO, K.P., doktor khim. nauk; TALMUD, S.L., kand. khim. nauk

Thermodynamics of woodpulp interaction with water and aqueous
solutions of electrolytes. Report No.5: Comparing the "active"
surfaces and heat of interaction with water of unbleached pulp
and viscose cellulose obtained therefrom before and after
fractionization. Trudy LTITSBP no.10:57-64 '62.

(MIRA 16:8)

(Woodpulp) (Heat of wetting) (Electrolytes)

FLIS, I.Ye., doktor khimicheskikh nauk; MISHCHENKO, K.P., doktor khimicheskikh nauk; TIMANOVA, T.A., kand.khimicheskikh nauk

Thermochemical study of the reduction reaction of chlorine dioxide and chlorine with sulfuric anhydride in water solutions at various temperatures. Trudy LTITSBP no.11:94-98 1962. (MIRA 1961)

PUSENOK, G.I., inzh.; FLIS, I.Ye., doktor khim.nauk; MISHCHENKO, K.P.,
doktor khim. nauk; BYNYAYEVA, M.K., kand.khim. nauk

Spectrophotometric method for studying the equilibrium of the
dissociation of hypobromous acid in aqueous solutions. Trudy
LTITSBP no.11:118-123 '62. (MIRA 14:10)

ARKHIPOVA, G.P., inzh.; ELIS, I.Ye., doktor khim.nauk; MISHCHENKO, I.I.,
doktor khim.nauk

Thermochemical study of the reduction of potassium chlorate by
sulfite in an acid medium. Trudy LTITSBP no.11:124-127 '62.

Spectrophotometric analysis of acid sulfite solutions. 128-133
(MIRA 1962)

FLIS, I.Ye.; MISHCHENKO, K.P.; SALNIS, K.Yu.

Study of the equilibrium $\text{ClO}_3 + \text{Cl}^- \rightleftharpoons 2\text{HClO}_2 + \text{O}_2 + \text{Cl}_2 + \text{H}_2\text{O}$ at
various temperatures. Zhur.prikl.khim. 34: 1011-1014, 1962.
(MIRA 1962)

(Chlorine oxides) (Phase rule and equilibrium)

KUSTODINA, V.A.; MISHCHENKO, K.P.; FLIS, I.Ye.

Thermodynamics of formation of chlorine monoxide in carbon
tetrachloride. Zhur.prikl.khim. 35 no.1:1374-1376 Je '62.
(MIRA 15:7)

(Chlorine oxides) (Carbon tetrachloride)
(Heat of formation)

MISHCHENKO, K.P.; POLTORATSKIY, G.M.

Pressure of benzene vapors at the range of temperatures from
11 to 40°C. Zhur.prikl.khim. 35 no.7:1638-1640 J1 '62.
(MIRA 15:8)

(Benzene)

(Vapor pressure)

MAKOLKIN, Ivan Afanas'yevich; SHMELEV, Boris Aleksandrovich;
IZMAYLOV, A.V., doktor khim. nauk, retsenzent;
KARAPET'YANTS, M.Kh., doktor khim. nauk, retsenzent;
MISHCHENKO, K.P., doktor khim. nauk, retsenzent;
FEDOROVA, T.P., red.; BARANOV, Yu.V., tekhn. red.

[Collection of examples and problems in physical and col-
loid chemistry] Sbornik primerov i zadach po fizicheskoi
i kolloidnoi khimii. Moskva, Rosvuzizdat, 1963. 181 p.
(MIRA 16:4)

(Chemistry, Physical--Problems, exercises, etc.)

PARAM, A.A.; KOKUSHKIN, O.A.; MISHCHENKO, K.P.; FLIS, I.Ye.

Laboratory study of the extraction of a complex catalyst from
polyethylene dispersions by methanol in rotary apparatus.
Plast. massy no.8:7-11 '63. (MIRA 16:8)

(Polyethylene) (Catalysts)

BARAM, A.A.; KOKUSHKIN, O.A.; MISHCHENKO, K.P.; FLIS, I.Ye.; ARKHIPOVA,
Z.V.; VAVILOVA, I.I.; MONAKHOVA, Ye.V.; SHCHUTSKIY, S.V.

Recovery of complex catalysts from dispersions of polyethylene
by means of methanol in a rotary apparatus. Plant. massy
no.11:58-59 '63. (MIRA 10:12)

MISHCHENKO, K.P.; SOKOLOV, V.V.

Thermodynamics and structure of nonaqueous solutions of electrolytes.
Part 3: Comparison of the structures of solutions of sodium
perchlorate in acetone and water. Zhur.strukt.khim. 4 no.2:184-188
Mr-Ap '63. (MIRA 16:5)

1. Leningradskiy tekhnologicheskiy institut tsellyulozno-bumazhnoy
promyshlennosti.

(Sodium perchlorate) (Acetone)
(Electrolyte solutions—Thermodynamic properties)

SOKOLOV, V.V., ZHILINA, L.P., MISHCHENKO, K.M.

Thermodynamics of the vaporization of acetone at various
temperatures. Zh. prikl. khim., 1964, 37, 78, Apr. 1964.

(Acetone) evaporation

OSINSKA-TANEVSKA, S.M.; BYNYAYEVA, M.K.; MISHCHENKO, K.P.; FLIS, I.Ye.

Spectrophotometric determination of the constants of dissociation
of hypochlorous acid at various ~~temperatures~~. Zhur.prikl.khim.
36 no.1212-1217 Jo '63. (MIRA 16:8)
(Hypochlorous acid) (Dissociation) (Spectrophotometry)

1. The first part of the document is a list of the names of the persons who were present at the meeting.

2. The second part of the document is a list of the names of the persons who were present at the meeting.

ABSTRACT: This report, K-1, is a summary of the results of the study.

The study was conducted in the laboratory of the Department of the Army, Fort Belvoir, Illinois, and was supported by the Army Research Office (ARO), Fort Belvoir, Illinois. The study was conducted in the laboratory of the Department of the Army, Fort Belvoir, Illinois, and was supported by the Army Research Office (ARO), Fort Belvoir, Illinois.

MISHCHENKO, K.I.; SIBIRSKY, V.V.

Thermodynamics and structure of nonelectrolyte solutions.
Part 5: Solutions of sodium iodide and perchlorate in acetone and
water at various temperatures. Zhur. strukt. khim. 5: 10, 1964, 222.
N-D 164. PPA 11:

1. Leningradskiy tekhnologicheskii institut tsellyulozno-bumazhnoy
promyshlennosti.

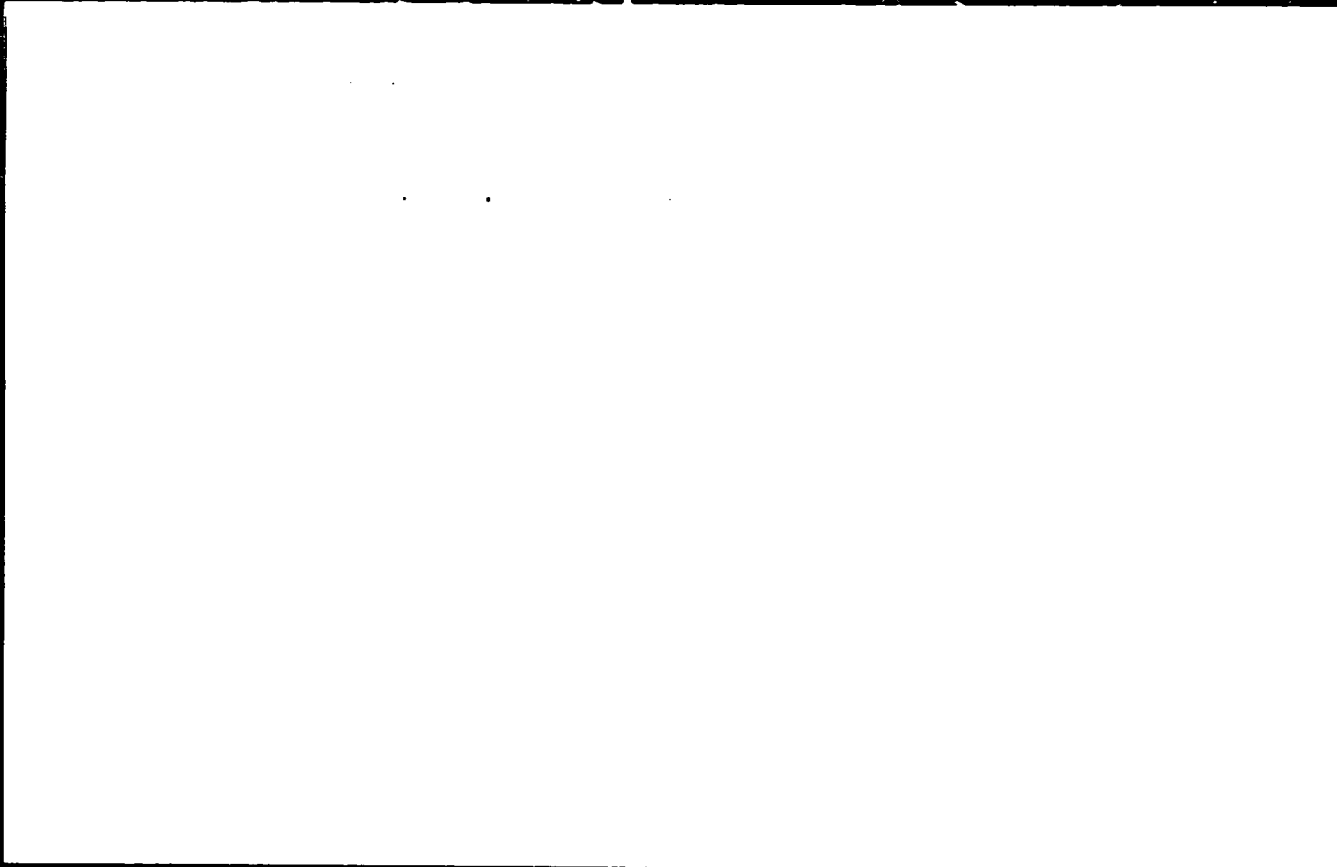
FLIS, I. Ye. [deceased]; MISHCHENKO, K.P.; PUSENOK, G.I.

Spectrophotometric determination of the dissociation constant of hypobromous acid at various temperatures. Izv. vys. ucheb. zav.; khim. i khim. tekhn. 7 no.5:764-767 '64. (MIRA 18:1)

1. Kafedra fizicheskoy i kolloidnoy khimii Leningradskogo tekhnologicheskogo instituta tsellyulozno-bumazhnoy promyshlennosti.

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L 38582-65

ACCESSION NR: AP5011046

UR/0080/64/037/010/2263/2268

AUTHOR: Dymarchuk, N. P.; Mishchenko, K. P.; Fomina, T. V.

TITLE: Characteristics of the molecular weight of cellulose triacetate obtained by acetylation of wood cellulose by a combination of osmometric and viscometric methods

SOURCE: Zhurnal prikladnoy khimii, v. 37, no. 10, 1964, 2263-2268

TOPIC TAGS: cellulosic plastic, cellulose, molecular weight

Abstract: Physical nonuniformity of cellulose triacetate obtained by acetylation of wood cellulose is studied in this report. A batch of partially saponified cellulose triacetate with an acetyl number of 61 and a degree of polymerization of the wood cellulose specimens equal to 1280 was selected for the investigation. To determine the chemical nonuniformity in the original samples and in subsequent fractions of the original sample, the content of bound acetic acid was determined. It was found that in all cases the content was 61%, that is, the fractions were chemically homogeneous. The molecular weight was calculated from data of viscosimetric measurements. Orig. art. has 6 formulas, 4 graphs, and 3 tables.

Card 1/2

L 38582-65

ACCESSION NR: AP5011046

ASSOCIATION: Leningradskiy tekhnologicheskii institut tsellyulozno-bumazhnoy
promyshlennosti (Leningrad Technological Institute of the Cellulose-Paper Industry)

SUBMITTED: 03Nov62

ENCL: 00

SUB CODE: MT, GC

NO REF SOV: 005

OTHER: 001

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SHALFIV, C.T.; MICHAKIS, A.T.

the 1950s, the number of people in the United States who were of South or Central American descent, particularly Mexican-Americans, increased. This was due to a variety of factors, including immigration and the growth of the Mexican-American population in the United States.

1. Leishmaniasis - caused by Leishmania protozoa. It is a parasitic disease that can affect various parts of the body, including the skin, mucous membranes, and internal organs. It is transmitted by sandflies.

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DANILOV, S.N., glavn. red.; ARBUZOV, A.Ye., red.; VVEDENSKIN, A.A., red.; VENUS-DANILOVA, E.L., red.; GABRILOVA, A.I., red.; IGARTE, I.S., red.; KALECHNEVA, Ye.L., red.; LONCHENKO, I.P., red.; MISHCHENKO, K.I., red.; NEFODI, M.G., red.; PETROV, A.A., red.; FREYDLINA, L.Kh., red.; SHEVAKIN, M.M., red.; SHUKAREV, S.A., red.; YURAYEV, Yu.K., red.

[Biologically active compounds] biologicheski aktivnye soedineniya. Moskva, Nauka, 1988. 306 s.

(NIA 1817)

MISHCHENKO, K.P.; TUNGUSOV, V.P.

Thermochemistry of nonaqueous electrolyte solutions. Part 1: Heats of dissolution in ethylene glycol of NaI at 25 and 25⁰ and of KI at 25⁰. Teoret. i eksper. khim. 1 no.1:55-59 Ja-F '65. (MIRA 18:7)

1. Leningradskiy tekhnologicheskij institut tsellyulozno-bumazhnoy promyshlennosti.

MISHCHENKO, K.P.; SHADSKIY, S.V.

Thermochemistry of nonaqueous electrolyte solutions. Part 2: Heats of dissolution of sodium iodide in a water-dioxane solvent. Teoret. i eksper. khim. 1 no.1:60-65 Ja-F '65. (MIRA 18:7)

1. Leningradskiy tekhnologicheskii institut tsellyulozno-bumazhnoy promyshlennosti.

DANILOV, S.N., red.; ZAKHAROVA, A.I., red.; ARBUZOV, A.Ye.,
red.; YVERLOVA, A.A., red.; YEMUS DANILOVA, E.T., red.;
LOFFE, I.S., red.; KAVYAZHEVA, Ye.I., red.; LUTSENKO,
I.F., red.; LUKHIN, K.F., red.; NEMTSEV, I.I., red.;
PETROV, A.A., red.; PRUTIN, R.Kh., red.; SHEYAKIN,
I.M., red.; SHUKAREV, S.A., red.; YER'YEV, Ye.K., red.

[Incluziune în volumul de referință al bibliotecii de la
Sinteză. Documentul este în limba rusă. Data 1980]

MISHCHENKO, P. I. PLYUSOVA, M. I.

Thermochemistry of nonequivalent electrolyte solutions. Part 3: heats
of solution of NaI and KI in methanol. Teoret. i eksper. khim. :
no. 2: 201-204. Mr-Apr 1965. (MIRA 12:17)

L. Leningradskiy tekhnicheskiy institut tsellyulozno-bumazhnoy
promyshlennosti.

ZHILINA, L.P.; MISHCHENKO, K.P.

Thermochemistry of nonaqueous electrolyte solutions. Part 4
Integral heats of dissolution of picric acid in acetone and
sodium picrate in methanol as a function of concentration at
10 and 25°C. Teoret. i eksper. khim. 1 no. 3: 361-366. My-Je '65.
(MIRA 18-9)

1. Leningradskiy tekhnologicheskij Institut tsellyulozno-
bunazhnoy promyshlennosti.

ТАНЕВСКА ОСИПКА Д.М. (Тане Мана Осипско, С.), станбен, М.Б.Т.Н.А.К.И., к.ф.

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